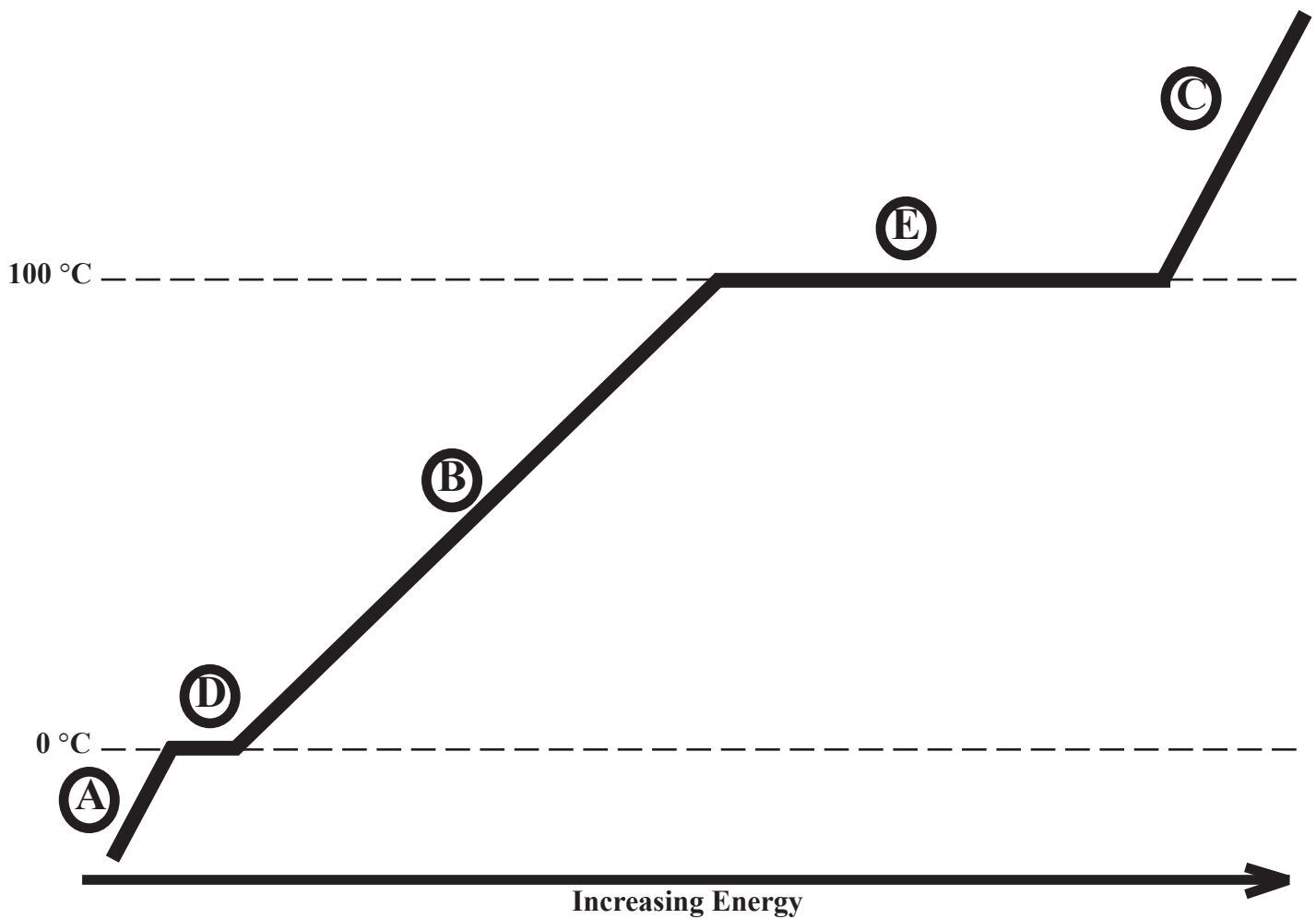


Changes of State - Water Std 7c



For the slopes – Specific Heat

(A) $S_{\text{(solid) Ice}} = 2.0 \text{ J/g}^\circ\text{C}$

(B) $S_{\text{(liquid) Water}} = 4.18 \text{ J/g}^\circ\text{C}$

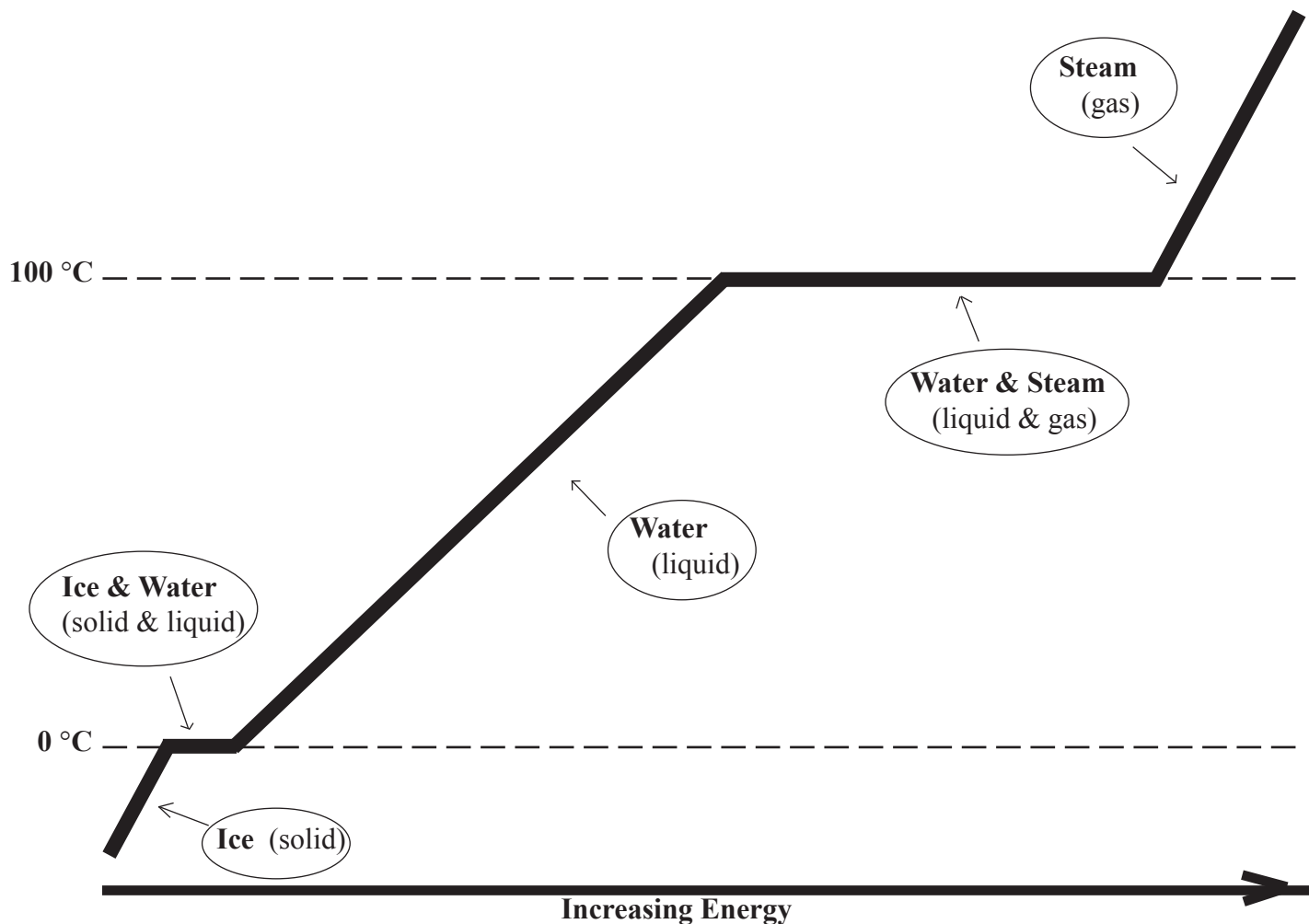
(C) $S_{\text{(gas) Steam}} = 2.0 \text{ J/g}^\circ\text{C}$

For Flat Places – Heat of . . .

(D) heat of **fusion** = 6.02 kJ/mole
(solid \leftrightarrow liquid)
ice \leftrightarrow water

(E) heat of **vaporization** = 40.6 kJ/mole
(liquid \leftrightarrow gas)
water \leftrightarrow steam

Changes of State - Water Std 7c



For the slopes – Specific Heat

(A) $S_{\text{(solid)}}$ = 2.0 J/g°C
Ice

(B) $S_{\text{(liquid)}}$ = 4.18 J/g°C
Water

(C) $S_{\text{(gas)}}$ = 2.0 J/g°C
Steam

For Flat Places – Heat of . . .

(D) heat of fusion = 6.02 kJ/mole
(solid \leftrightarrow liquid)
ice \leftrightarrow water

(E) heat of vaporization = 40.6 kJ/mole
(liquid \leftrightarrow gas)
water \leftrightarrow steam