

# Memorize Ions Project - HELP

(MISC)

<u>Monatomic Ions</u>		<u>Polyatomic Ions</u>		<u>Acids</u>	
H <sup>1+</sup>	hydrogen ion	C <sub>2</sub> H <sub>3</sub> O <sub>2</sub> <sup>1-</sup>	acetate ion	HC <sub>2</sub> H <sub>3</sub> O <sub>2</sub>	acetic acid
Al <sup>3+</sup>	aluminum ion	NH <sub>4</sub> <sup>1+</sup>	ammonium ion	HCl	hydrochloric acid
At <sup>1-</sup>	astatide ion	CO <sub>3</sub> <sup>2-</sup>	carbonate ion	HF	hydrofluoric acid
B <sup>3+</sup>	boron ion	ClO <sub>3</sub> <sup>1-</sup>	chlorate ion	HNO <sub>3</sub>	nitric acid
Ba <sup>2+</sup>	barium ion	ClO <sub>2</sub> <sup>1-</sup>	chlorite ion	H <sub>3</sub> PO <sub>4</sub>	phosphoric acid
Be <sup>2+</sup>	beryllium ion	CrO <sub>4</sub> <sup>2-</sup>	chromate ion	H <sub>2</sub> SO <sub>4</sub>	sulfuric acid
Br <sup>1-</sup>	bromide ion	Cr <sub>2</sub> O <sub>7</sub> <sup>2-</sup>	dichromate ion		
Ca <sup>2+</sup>	calcium ion	OH <sup>-</sup>	hydroxide ion		
Cl <sup>1-</sup>	chloride ion	ClO <sup>1-</sup>	hypochlorite ion		
Cs <sup>1+</sup>	cesium ion	NO <sub>3</sub> <sup>1-</sup>	nitrate ion		
F <sup>1-</sup>	fluoride ion	ClO <sub>4</sub> <sup>1-</sup>	perchlorate ion		
Ga <sup>3+</sup>	gallium ion	PO <sub>4</sub> <sup>3-</sup>	phosphate ion		
In <sup>3+</sup>	Indium ion	SO <sub>4</sub> <sup>2-</sup>	sulfate ion		
I <sup>1-</sup>	iodide ion				
K <sup>1+</sup>	potassium ion				
Li <sup>1+</sup>	lithium ion				
Mg <sup>2+</sup>	magnesium ion				
Na <sup>1+</sup>	sodium ion				
O <sup>2-</sup>	oxide ion				
Po <sup>2-</sup>	polonide ion				
Rb <sup>1+</sup>	rubidium ion				
Se <sup>2-</sup>	selenide ion				
Sr <sup>2+</sup>	strontium ion				
Te <sup>2-</sup>	telluride ion				
Tl <sup>3+</sup>	thallium ion				

remember - the charges on all the monatomic ions can be figured by what column they are in

For groups 6 (16) and 7 (17) change the ending to -ide when the atom forms an ion.  
For groups 1, 2 & 3 (13) use the name of the metal for the name of the ion.