

Chapter 10 topics

Abbreviations for reaction types

- a. SYN - Synthesis – p. 284 \longrightarrow $A + B \longrightarrow C$
- b. COM - Combustion – p. 285 \longrightarrow $A + O_2 \longrightarrow C$
- c. DEC - Decomposition – p. 286 \longrightarrow $A \longrightarrow B + C$
- d. SR - Single Replacement – p. 287 \longrightarrow $AB + C \longrightarrow CB + A$
- e. DR - Double Replacement - p. 290 \longrightarrow $AB + CD \longrightarrow AD + CB$
- f. Ox-Red - Oxidation/Reduction (aka redox) – p. 636
- g. AB – Acid/Base - p. 295 \longrightarrow $\text{Acid} + \text{Base} \longrightarrow \text{H}_2\text{O} + \text{salt}$ Charge on ions chg

Acids and bases

Salt = metal + non-metal \dashrightarrow Always ionic

Acids have H⁺ ions \dashrightarrow on the front end

Bases have OH⁻ ions \dashrightarrow on the back end

Acid Examples HCl H₂SO₄ H₃PO₄

Base Examples: NaOH Ca(OH)₂ Al(OH)₃



always think of water as HOH